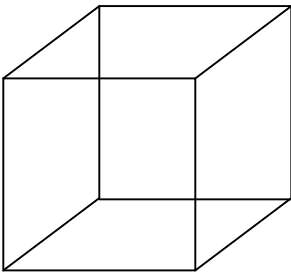
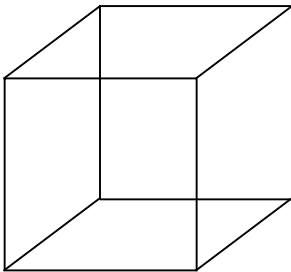
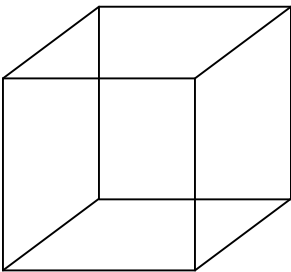


## Solid State Quick Reference Sheet

To aid in your understanding and studying of the material on solid state, fill in this table with the information you learn during the lab and keep this for a quick reference.

|   | Simple Cubic (sc)   | Body-Centered Cubic (bcc)  | Face-Centered Cubic (fcc)   |
|---|---|--|---|
| Draw the Image                            |  |  |  |
| Number of atoms in a unit cell            |   |  |   |
| Edge length (in terms of r)               |   |  |   |
| Volume of unit cell (in terms of r)       |   |  |   |
| Packing efficiency                        |   |  |   |
| Coordination Number of atoms in unit cell |   |  |   |
| Stacking pattern                          |   |  | ***   |

\*\*\*NOTE: To correctly determine the stacking pattern for the simple cubic and the body-centered cubic, you can look at the layers like you would for a z-diagram. However, for face-centered cubic, this is not the correct way. For face-centered cubic (since it is hexagonal packing while the others are square packing), you need to look along the diagonal THROUGH the entire unit cell to determine the correct stacking pattern (an example of the line you need to look along is illustrated on this cube below).

